

# Download Balanced Chemical Equation For The Combustion Of Octane

Ethanol (C<sub>2</sub>H<sub>5</sub>OH) is composed of C, H and O. So, the usual products of its combustion should be CO<sub>2</sub> and H<sub>2</sub>O. The unbalanced equation is: C<sub>2</sub>H<sub>5</sub>OH + O<sub>2</sub> = CO<sub>2</sub> + H<sub>2</sub>O Since both C and H are appearing at only once place each side, they should be balanced ...Methanol's chemical formula is CH<sub>3</sub>OH so the basic equation to burn this fuel in oxygen would be: nCH<sub>3</sub>OH + X O<sub>2</sub> = n CO<sub>2</sub> + 2n H<sub>2</sub>O. where n is the number of moles of methanol and X is the number of moles of oxygen required for complete combustion of methanol. Using Gibbs Free Energy (G) to Determine if a Chemical Reaction is Spontaneous or Nonspontaneous. According to the Second Law of Thermodynamics, the entropy of the universe increases, that is, the total entropy (ΔS total) increases. 8. Assume that 13.5 grams of solid aluminum react with HCl according to the following balanced equation at STP: 2 Al (s) + 6 HCl (aq) → 2 AlCl<sub>3</sub> (aq) + 3 H<sub>2</sub> (g)